



Improvement in photocatalytic activity of stable WO_3 nanoplatelet globular clusters arranged in a tree-like fashion: Influence of rotation velocity during anodization

R.M. Fernández-Domene, R. Sánchez-Tovar, B. Lucas-Granados, J. García-Antón*

Ingeniería Electroquímica y Corrosión (IEC), Departamento de Ingeniería Química y Nuclear, ETSI Industriales, Universitat Politècnica de València, Camino de Vera s/n, 46022 Valencia, Spain

ARTICLE INFO

Article history:

Received 15 January 2016

Received in revised form 25 February 2016

Accepted 27 February 2016

Available online 2 March 2016

Keywords:

WO_3 nanostructures

Anodization

Hydrodynamic conditions

Photocatalytic activity

Water splitting

ABSTRACT

This study investigates the influence of controlled hydrodynamic conditions during anodization of tungsten (W) on the morphological, electrochemical and photocatalytic properties of a novel WO_3 nanostructure: globular clusters of nanoplatelets associated in a tree-like fashion. For this purpose different techniques such as Field-Emission Scanning Electronic Microscopy (FE-SEM), electrochemical impedance spectroscopy (EIS) measurements, Mott-Schottky (M-S) analysis and photoelectrochemical water splitting tests have been carried out. Photoanodes obtained at 375 rpm showed the best photoresponse, much higher than that of conventional WO_3 nanoplatelets, which can be ascribed to a noteworthy increase in the electrochemically active surface area leading to improved charge transfer at the interface. Moreover, the improved WO_3 nanostructure displayed very good long-term photostability when irradiated with AM 1.5 illumination, thus solving the recurrent problem of the poor stability of WO_3 against photodegradation processes.

© 2016 Elsevier B.V. All rights reserved.

1. Introduction

Semiconducting metal oxides, such as TiO_2 [1–3], WO_3 [4–10], Fe_2O_3 [11,12] or ZnO [13,14], are being widely investigated for their use as photoanodes due to the wide range of properties they can offer. Among the key functional properties that these oxides must have for photoelectrochemical processes, the resistance to degradation in the required environments (chemical dissolution, electrochemical corrosion, photocorrosion and other light-driven degradation processes) is of critical importance [15,16]. Therefore, when using a photoelectrode, it is essential to assure that it is resistant to degradation and that its photoelectrochemical response is stable with time.

On the other hand, tungsten trioxide (WO_3) has generated great interest in the last years due to its numerous applications in electrochromic devices, gas sensors, photoelectrochemical water splitting or photodegradation of organic compounds [6–8,10,17–27]. To increase the efficiency of the WO_3 -

based photoanodes, nanostructured electrodes (such as nanotubes, nanorods/nanowires, nanoplatelets or nanoflakes) are required to provide a high specific surface area [4–10,18–20,22–31]. Although WO_3 is resistant to chemical attack in acidic media [26,32–34], its stability under illumination is not good in numerous electrolytes, even in strong acid solutions, which has been manifested in many works as a decay of the initial photocurrent density with time upon illumination [30,35–38]. In a recent paper [39] we presented a novel WO_3 nanostructure with very high surface area formed by tree-like nanoplatelets clusters. This innovative nanostructure was synthesized by anodization in sulfuric acid media in the presence of fluoride anions and under controlled hydrodynamic conditions using a rotating disc electrode (RDE), and its photoresponse was very high compared with that of samples synthesized under stagnant conditions. Although the new nanostructure formed under hydrodynamic conditions presented larger surface areas and provided higher photocurrent densities than that formed at 0 rpm, it was also observed that at high rotation velocities (2000 rpm) the thickness of the nanoplatelet layer, and hence the surface area exposed to the electrolyte, was significantly lower than at lower rotation velocities (500 rpm). This fact indicated that the influence of hydrodynamic conditions on the formation and growth of the

* Corresponding author.

E-mail address: jgarcia@iqn.upv.es (J. García-Antón).

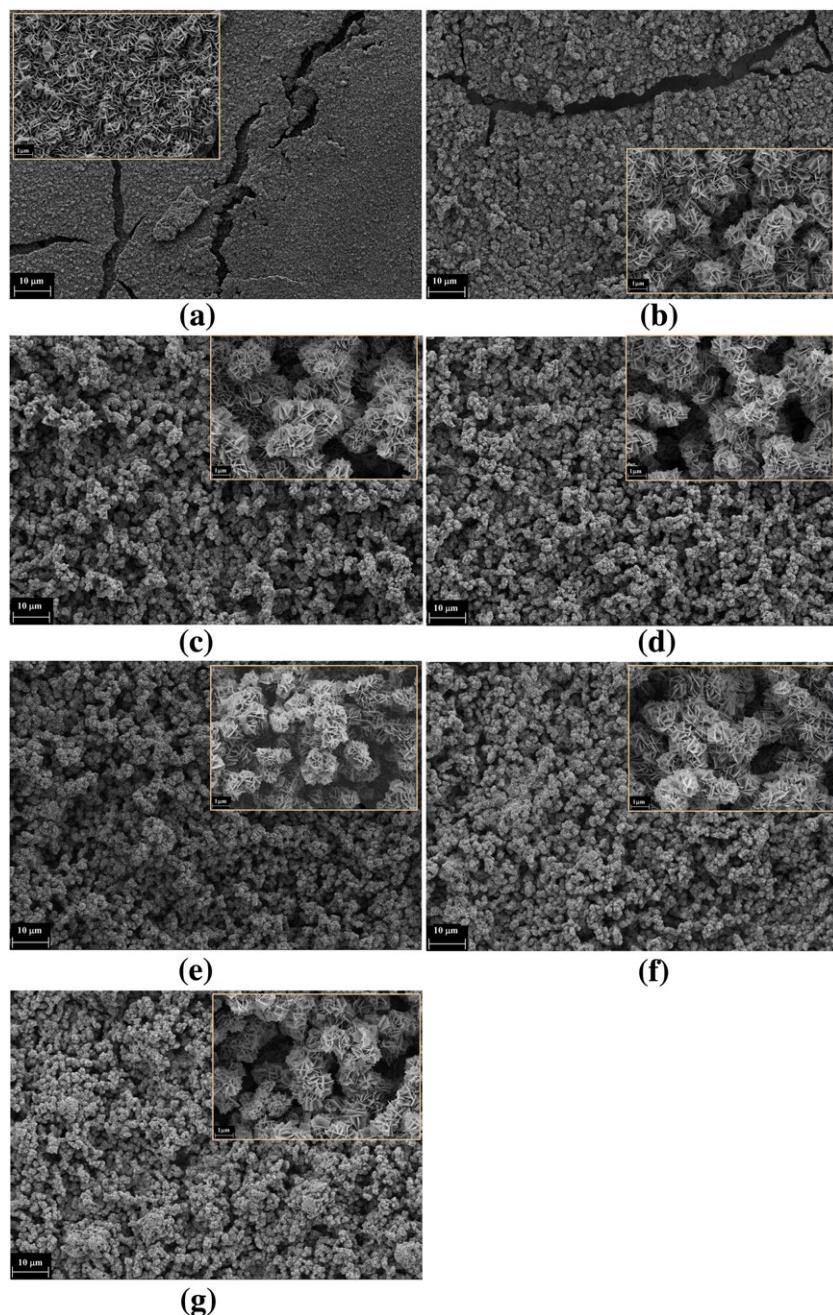


Fig. 1. FE-SEM images of the WO_3 nanostructures formed upon anodization of W in a $1.5 \text{ M H}_2\text{SO}_4 + 0.1 \text{ M NaF}$ electrolyte at 50°C , imposing a potential of 20 V for 4 h , under different hydrodynamic conditions: (a) 0 rpm , (b) 125 rpm , (c) 250 rpm , (d) 375 rpm , (e) 500 rpm , (f) 750 rpm , and (g) 1000 rpm .

new tree-like nanostructure was not linear, and further study was necessary to investigate the mechanism of formation of these WO_3 nanoplatelets clusters.

The objectives of the present study are, therefore: (1) to optimize the dimensions and photoactivity of the tree-like nanoplatelets by modifying the rotation velocity of the RDE and (2) to find optimal conditions to assure that the new nanostructure does not undergo degradation under illumination, that is, its photoresponse remains stable with time. To accomplish both objectives, different techniques have been used, such as Field-Emission Scanning Electronic Microscopy (FE-SEM), electrochemical impedance spectroscopy (EIS), Mott-Schottky (M-S) analysis and photoelectrochemical measurements.

2. Experimental procedure

Tungsten samples were teflon-coated cylindrical bars of 8 mm in diameter. Before anodization, samples were abraded using 220, 500, 1000 and 4000 grit SiC papers to remove scratches and other imperfections from the surface. Then, the samples were sonicated in ethanol for 2 min and dried in a N_2 stream. Anodization of tungsten samples was carried out in a $1.5 \text{ M H}_2\text{SO}_4 + 0.1 \text{ M NaF}$ electrolyte, at 20 V and 50°C for 4 h , under controlled hydrodynamic conditions (from 0 to 1000 rpm) by using a rotating disc electrode (RDE) and a platinum mesh as cathode. The area of the working electrode (tungsten sample) exposed to the electrolyte was 0.5 cm^2 . After anodization, samples were annealed at 400°C for 4 h in the presence of oxygen. The morphology of the formed nanostructures

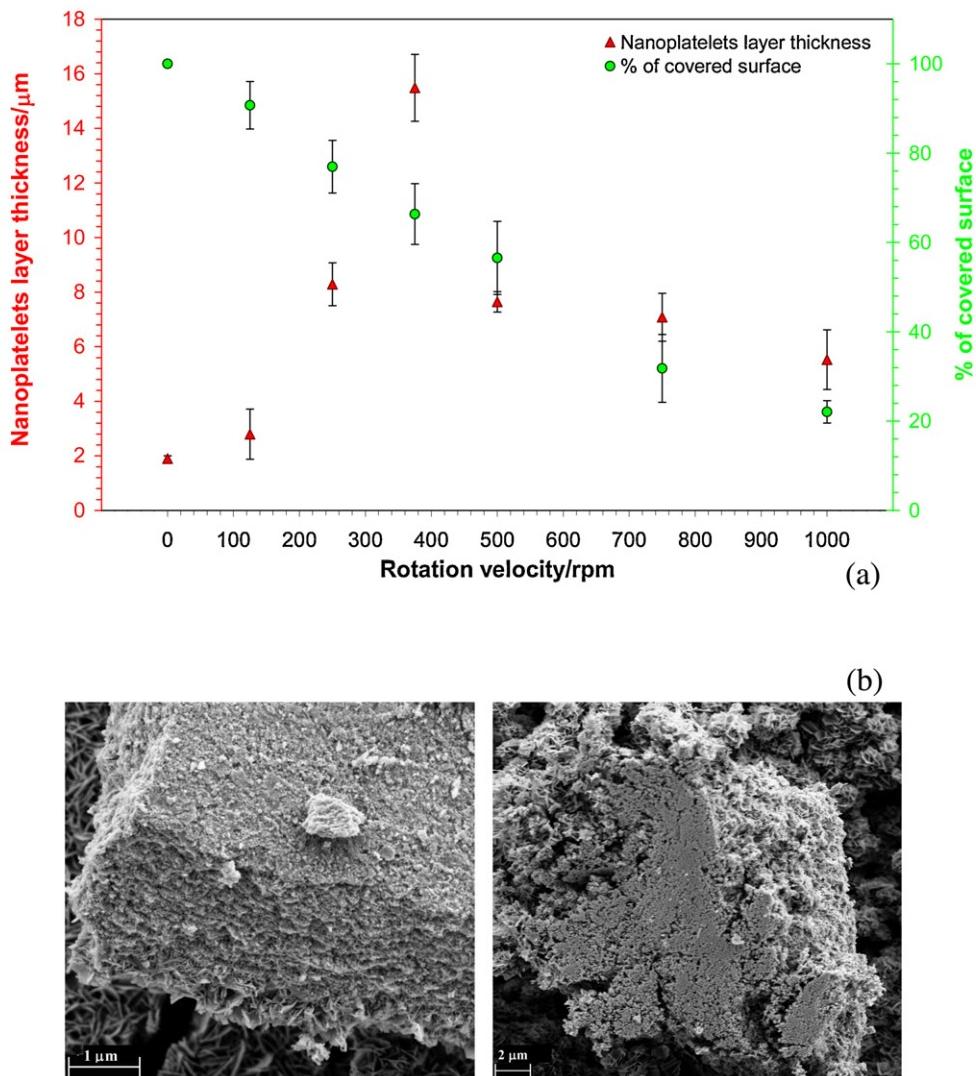


Fig. 2. Nanoplatelets layer thickness and% of covered surface as a function of the rotation velocity (a). FE-SEM cross-sectional images of samples anodized at 0 rpm (left) and 500 rpm (right).

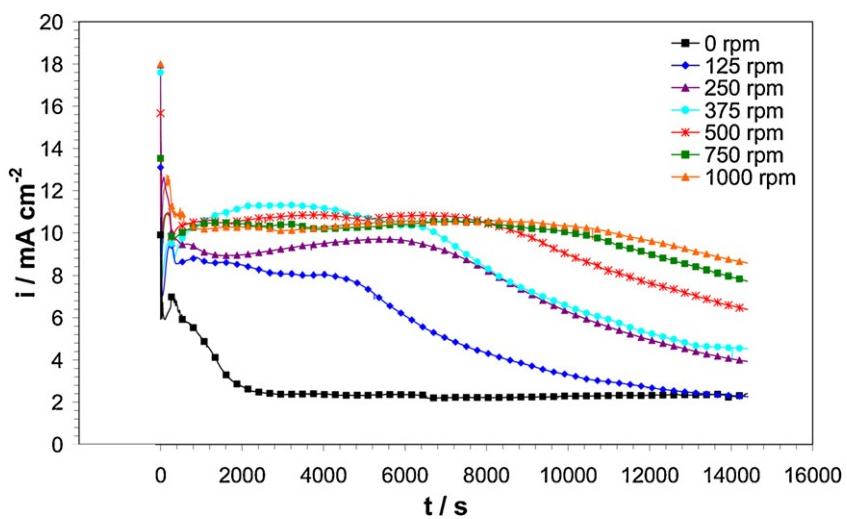


Fig. 3. Current density transient recorded during anodization for all the samples.

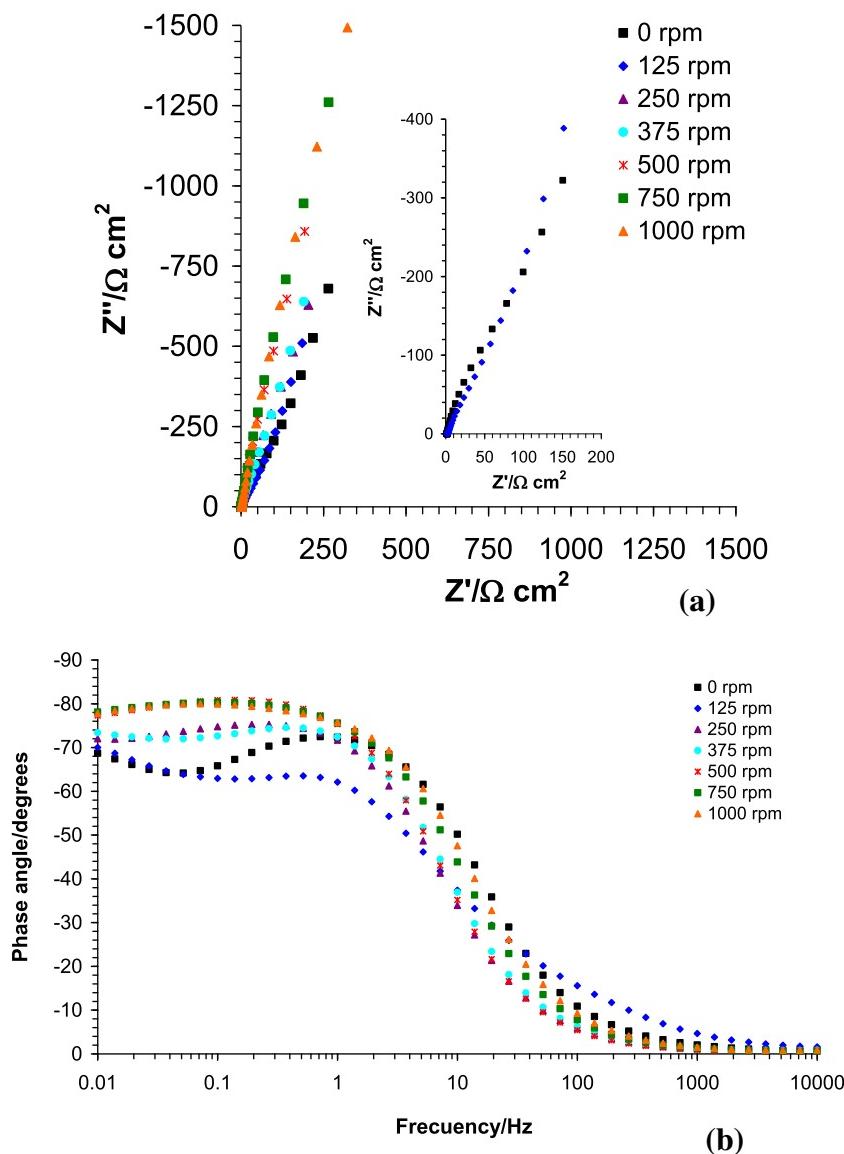


Fig. 4. Experimental Nyquist (a) and Bode-phase (b) diagrams for the WO_3 nanplatelets formed at different rotation velocities.

was characterized by Field Emission Scanning Electron Microscopy (FE-SEM).

Electrochemical measurements (electrochemical impedance spectroscopy (EIS) and Mott-Schottky analysis) were performed under dark conditions in a 1 M HCl electrolyte. An Ag/AgCl (3 M KCl) electrode was the reference electrode, and a platinum tip was the counter electrode. The illuminated area of the working electrode exposed to the electrolyte was 0.13 cm^2 . For EIS measurements at open circuit potential, the potential perturbation amplitude was 10 mV, in the frequency range of 10 kHz to 10 mHz. Mott-Schottky plots were determined at a constant frequency of 5 kHz using a 10 mV amplitude signal and scanning the potential from a value of $0.5 \text{ V}_{\text{Ag}/\text{AgCl}}$ in the negative direction at a rate of 50 mV s^{-1} .

Photocurrent density-potential curves and stability tests were carried out in many different electrolytes using an Autolab PGSTAT302 N potentiostat under simulated sunlight illumination AM 1.5 (100 mW cm^{-2}), with the purpose of finding an electrolyte where the WO_3 samples were stable. The electrochemical cell was the same as that described above for the electrochemical measurements. Photocurrent density-potential curves were recorded by scanning the potential from -0.24 V to 1.02 V with a scan rate

of 2 mV s^{-1} , switching the light on and off (60 s in the dark and 20 s in the light) to obtain in the same measurement photocurrent and dark current values. Afterwards, a potential of 1.02 V was kept constant for 1 h under illumination to evaluate the samples stability. Additionally, the indirect band-gap energy of the samples was determined measuring the photocurrent density spectrum at $1.02 \text{ V}_{\text{Ag}/\text{AgCl}}$ in the wavelength range from 300 to 500 nm, in a 1 M HCl electrolyte.

3. Results and discussion

3.1. Morphological characterization (FE-SEM)

Fig. 1 shows the FE-SEM images of the samples anodized in the 1.5 M H_2SO_4 /0.1 M NaF solution at different rotation velocities. It can be observed that in all cases, the nanostructures formed on the W surface presented nanplatelet morphology, as other authors reported when anodizing in similar acid electrolytes [18,20,25,29,39,40]. However, the dimensions and morphology of these nanplatelets were influenced by hydrodynamic conditions. At 0 rpm (**Fig. 1a**), nanplatelets grew quite orderly and almost

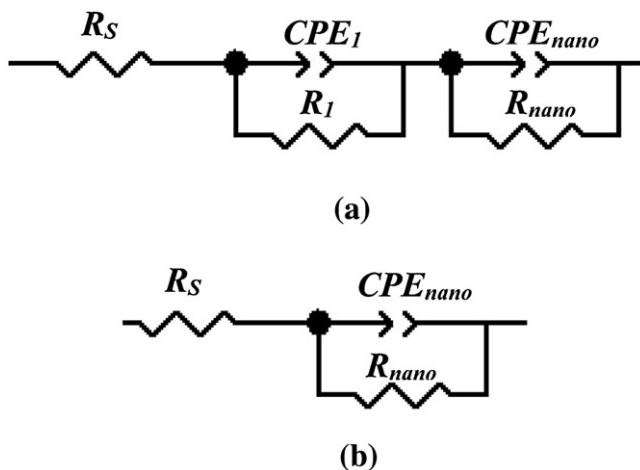


Fig. 5. Electrical equivalent circuits used to simulate the experimental EIS data for the WO_3 nanplatelets formed at (a) 0, 125, 250 and 375 rpm; (b) 500, 750 and 1000 rpm.

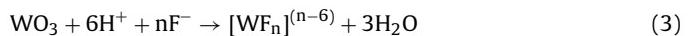
orthogonal to the electrode surface. The nanplatelets layer was crisscrossed by numerous cracks and at least another nanplatelets layer could be observed underneath the first one. At 125 rpm (Fig. 1b), nanplatelets started forming globular clusters, although most of these clusters still grew forming a single-level layer; in this case, several cracks were also observed on the WO_3 layer. At rotation velocities higher than 125 rpm (Fig. 1c–g), nanplatelets globular clusters grew in a tree-like manner, forming a multilevel WO_3 layer resembling a coniferous forest or a coral reef seen from above and thus exposing more surface area to the electrolyte. It can be observed that at high rotation velocities (Fig. 1e–g) these nanplatelet clusters were denser than at lower rpm, although their size did not vary appreciably with the rotation velocity. Fig. 2a plots the total thickness of the nanplatelets layer and the% of surface covered by these nanostructures as a function of the rotation velocity. The total thickness was obtained from FE-SEM cross-sectional images. Fig. 2b shows, as an example, the FE-SEM cross-sectional images of the samples anodized under 0 and 500 rpm. It can be seen that the thickness of the WO_3 nanplatelets layer increased with increasing rpm until reaching a maximum value at 375 rpm, and then decreased again. Concerning the% of surface coverage, it decreased more or less linearly with rotation velocity. These differences in dimensions and surface coverage will be discussed in the next section when studying the current density transients recorded during the anodization process.

3.2. Influence of rotation velocity on the anodization process

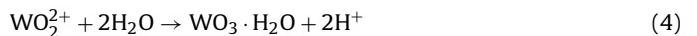
In order to understand the influence of rotation velocity on the formation and growth of WO_3 nanplatelets, current density transients recorded during the anodization of the W samples are presented in Fig. 3. A notable decrease in current density can be observed during the first seconds of anodization, regardless of the rotation velocity. This decrease is associated with the formation of a compact WO_3 layer on the electrode surface [4,39,41], according to:



Subsequently, current density started increasing, indicating the dissolution of the WO_3 layer due to the effect of H^+ [24,29,39,42] and F^- [24,39,41]. This dissolution process resulted in the release of soluble species from the electrode surface:



Afterwards, current density gradually decreased again, eventually reaching a steady-state value. This decrease and the later stabilization of current density can be explained by the precipitation of soluble species upon reaching supersaturation conditions near the electrode surface. This precipitation in the form of nanplatelets took place according to the following reactions [24,29,39,42]:



In a previous work is was demonstrated through Raman spectroscopy that the main constituent of the hydrated WO_3 nanplatelets was $\text{WO}_3 \cdot 2\text{H}_2\text{O}$, although the presence of some $\text{WO}_3 \cdot \text{H}_2\text{O}$ was also detected [39]. After annealing the samples at 400 °C for 4 h, crystalline monoclinic WO_3 was obtained [39].

Under static conditions (0 rpm), current densities recorded during the whole anodization process were notably lower than under hydrodynamic conditions. Moreover, steady-state values were attained much earlier than with rotation, indicating that the WO_3 nanplatelet layer completely covered the electrode surface after ~2000 s from the beginning of the anodization process.

As the rotation velocity increased, current density values associated with the dissolution of the initial WO_3 layer also increased. This fact is explained by taking into account the velocity profile of the electrolyte near the RDE. When the RDE rotates, the solution is drawn up from the bulk electrolyte towards the electrode along the axis normal to the surface, and afterwards, due to the centrifugal force, it is expelled towards the outer edge of the electrode. The expression for the fluid velocity in the normal direction, v_y , is [43]:

$$v_y = -0.51\omega^{3/2}\nu^{-1/2}y^2 \quad (6)$$

where ω is the angular velocity, ν the kinematic viscosity of the fluid, and y (in cylindrical coordinates) the distance from the electrode surface in the normal direction.

Eq. (6) implies that v_y increases (in absolute value) with increasing the rotation velocity (ω). Hence, faster electrode rotation permitted the ions in the electrolyte (F^- and H^+) to transport to the electrode more rapidly (both by convection and by diffusion through the stagnant layer), resulting in higher current densities being measured at the working electrode due to the enhancement of the WO_3 dissolution (Eqs. (2) and (3)), and the subsequent exposure of the W bare surface to the electrolyte, which oxidized again (Eq. (1)). Consequently, an increase in the rotation velocity increased the release of soluble species from the electrode surface, which would result in a higher precipitation rate of hydrated WO_3 in the form of nanplatelets (Eqs. (4) and (5)) due to supersaturation conditions being reached earlier than at low ω . However, it was observed before from the FE-SEM images that the thickness of the WO_3 nanplatelets layer, as well as the degree of surface covered by them, were smaller at high rotation velocities (Fig. 2). This is explained by the existence of a radial flow because of the centrifugal force that threw the electrolyte outwards from the center of the electrode, with the fluid velocity increasing towards the edge of the disc with increasing distance from the rotation axis in the radial direction (r) and ω , according to [43]:

$$v_r = 0.51\omega^{3/2}\nu^{-1/2}ry \quad (7)$$

Consequently, soluble species near the electrode surface were constantly swept away radially at a velocity directly dependent on the rotation velocity and the radius. As described above, for nanplatelets to precipitate on the electrode surface, soluble species must reach a supersaturation concentration near the elec-

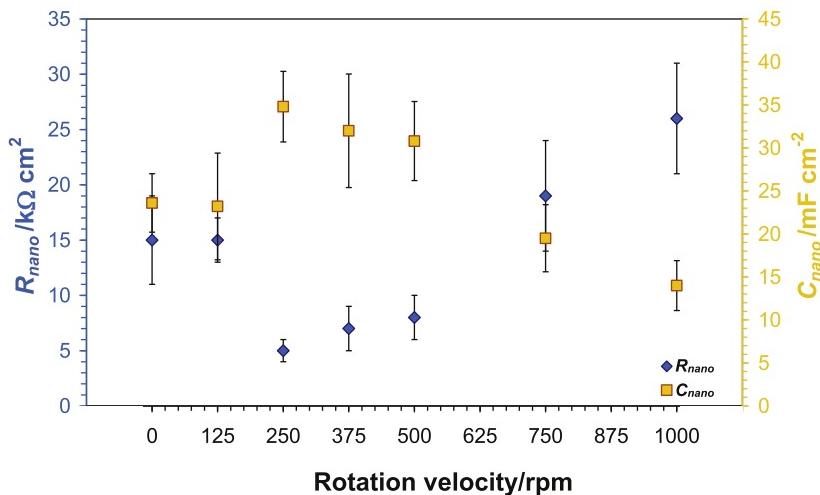


Fig. 6. Values of the charge-transfer resistance (R_{nano}) and the double layer capacitance (C_{nano}) of the WO_3 nanoplatelets formed at different rotation velocities.

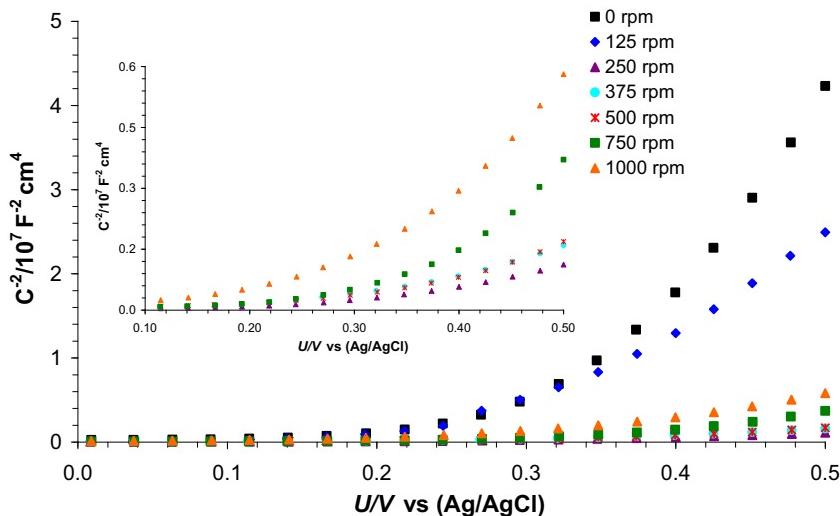


Fig. 7. Mott-Schottky plots for the WO_3 nanoplatelets formed at different rotation velocities.

trode. As ω increased, this critical concentration was more difficult to accomplish, especially at high values of r .

Therefore, hydrodynamic conditions influenced the growth of nanoplatelets in two opposite ways: on the one hand, hydrodynamic conditions enhanced mass transfer towards the electrode surface, thus favoring the formation of soluble species responsible for the formation of nanoplatelets; on the other hand, mechanical convection reduced the concentration of soluble species close to the electrode surface, hence decreasing the nanoplatelets growth rate by precipitation [5,29,39]. In this case, according to the FE-SEM images and the nanoplatelet layer thicknesses shown in Fig. 2, the first factor (enhancement of mass transfer, soluble species formation and subsequent precipitation) had more influence at low rotation velocities (up to 375 rpm), while at rotation velocities higher than 375 rpm it was the second factor (decrease of soluble species concentration near the surface) that predominated over the first one in determining the formation and growth of nanoplatelets, since the thickness of the nanoplatelets layer decreased at high rpm.

Current densities recorded at the end of the 4-hours anodization process also increased with increasing rotation velocity. These values represent an approximately steady-state between the dissolution of the WO_3 layer (Eq. (1)) and the precipitation of soluble

species to hydrated WO_3 in the form of nanoplatelets (Eqs. (4) and (5)). Hence, an increase in steady-state current density with ω implies that the degree of surface coverage was smaller at high ω , which is also consistent with the previous results (Fig. 2).

3.3. Electrochemical characterization

The electrochemical characterization of the WO_3 nanostructures/electrolyte interface was performed in the dark through EIS measurements and Mott-Schottky analysis.

Fig. 4 shows the Nyquist and Bode-phase diagrams for all the samples, anodized at different rotation velocities. Nyquist plots (Fig. 4a) exhibit one or two unfinished semicircles, depending on the rotation velocity, and relatively low impedance values. At 0 and 125 rpm, two semicircles can be discerned, one at intermediate frequencies and the other at low frequencies (inset in Fig. 4a), which correspond with the two peaks that appear in the Bode-phase plots (Fig. 4b). At 500 rpm and higher rotation velocities (750 and 1000 rpm), only one semicircle can be discerned from Nyquist plots, in accordance with the peak observed in the phase angle. At 250 and 375 rpm, a single semicircle is appreciable in Nyquist plots, but two time constants can be detected from the Bode-phase plots. The

first semicircle observed in the Nyquist plots at 0 and 125 rpm has a considerably lower amplitude than the second semicircle, being the amplitude of the latter more similar to that of the single semicircle observed at higher rotation velocities. Therefore, this first time constant can be associated with some surface effects related to the specific morphology of the samples anodized at low rotation velocities, especially at 0 and 125 rpm. At 0 rpm, the morphology of the WO_3 nanostructures was observed to be vertically-alligned nanoplatelets, while at 125 rpm this morphology slightly changed and single nanoplatelets started forming globular clusters, but in both cases nanoplatelets formed a single-level layer (Fig. 1). As rotation velocity increased, globular clusters of nanoplatelets began forming multilevel layers, and this first time constant was not so clearly observed from EIS plots (at 250 and 375 rpm, where a transition occurred) or simply disappeared (500, 750 and 1000 rpm). The second time constant, which is common for all the samples, can consequently be associated with the charge-transfer resistance and the double layer capacitance of the WO_3 nanostructures.

Since the impedance response of the samples, especially the variation of the phase angle with frequency, was observed to change depending on the rotation velocity used during anodization, EIS experimental data have been fitted to the electrical equivalent circuits shown in Fig. 5a (from 0 to 375 rpm) and in Fig. 5b (from 500 to 1000 rpm). In these circuits, R_S is the electrolyte resistance; the $R_1 - \text{CPE}_1$ time constant is associated with some surface effects due to the morphology of the nanoplatelet layer formed at 0 and low rpm values (mainly 125 rpm), while the $R_{\text{nano}} - \text{CPE}_{\text{nano}}$ time constant is related to the charge transfer between the WO_3 nanostructure and the electrolyte. Constant Phase Elements (CPEs) have been used instead of pure capacitors, to account for the non-ideality of the system under measurement. The impedance of a constant-phase element is defined as:

$$Q = Z_{\text{CPE}} = [C(j\omega)^n]^{-1} \quad (8)$$

where n , defined as a CPE power, is an adjustable parameter that lies between –1 and 1. For $n = 1$ the CPE describes an ideal capacitor, and for $n = 0$ the CPE is an ideal resistor. For $0.5 < n < 1$ the CPE describes a frequency dispersion of time constants due to local heterogeneities in the dielectric material. A pure inductance yields $n = -1$.

CPEs of the WO_3 nanostructures have been converted into pure capacitances, C , by using the following equation [44,45]:

$$C_{\text{nano}} = \frac{(Q_{\text{nano}} \cdot R)^{1/\alpha}}{R} \quad (9)$$

where Q_{nano} is the impedance of the CPE_{nano} . For the equivalent circuit shown in Fig. 5a, $R = R_{\text{nano}}$, while for the equivalent circuit in Fig. 5b, $R = R_S \cdot R_{\text{nano}}/R_S + R_{\text{nano}}$ [44,45].

The values of R_{nano} and C_{nano} have been plotted against the rotation velocity in Fig. 6. A clear correlation between both values is evident: high R_{nano} values imply low C_{nano} values, and vice versa. At 0 and 125 rpm, both parameters were similar, but as rotation velocity increased (up to 500 rpm), R_{nano} decreased and C_{nano} increased. This fact can be explained by an increase in the surface area exposed to the electrolyte, which increased the double layer capacitance and reduced the charge-transfer resistance at the nanostructure/electrolyte interface [41] due to a large number of electrochemically active sites being permeated by the electrolyte [8]. At higher rotation velocities, R_{nano} increased again whereas C_{nano} decreased. Although the morphology of the WO_3 nanostructures formed at 750 and 1000 rpm was similar than that obtained at lower rpm (Fig. 1), nanoplatelets clusters formed at high rotation velocities were observed to grow in a more compact manner and, besides, the thickness of the WO_3 nanoplatelet layer was lower for these samples (Fig. 2), which resulted in a decrease of the active surface area exposed to the electrolyte. These factors led to a reduction in the efficiency of electron transport [30] and, therefore, the

Table 1

Donor density (N_D) and flat-band potential (U_{FB}) of the WO_3 nanoplatelets formed at different rotation velocities.

Sample (rpm)	$N_D (\times 10^{22} \text{ cm}^{-3})$	$U_{FB}/V_{\text{Ag/AgCl}}$
0	1.2 ± 0.3	0.30 ± 0.07
125	2.4 ± 0.2	0.29 ± 0.03
250	49.5 ± 19.1	0.28 ± 0.06
375	41.9 ± 8.1	0.32 ± 0.07
500	57.3 ± 13.5	0.23 ± 0.01
750	14.3 ± 0.5	0.30 ± 0.01
1000	10.5 ± 0.1	0.23 ± 0.02

charge-transfer resistance increased and the double layer capacitance decreased at 750 and 1000 rpm.

In order to investigate the semiconducting properties of the WO_3 nanoplatelets layer, the electrochemical interfacial capacitance was measured as a function of the applied potential. Mott-Schottky (M-S) analysis was performed to obtain the donor density, N_D , and the flat-band potential, U_{FB} , of the nanostructures formed at different rotation velocities. Fig. 7 shows the Mott-Schottky plots of the different samples. The positive slopes of these M-S plots are characteristic of *n*-type semiconductors, and the nanostructure/electrolyte interface can be described by the corresponding Mott-Schottky equation:

$$\frac{1}{C^2} = \frac{1}{C_H^2} + \frac{2}{\varepsilon_r \varepsilon_0 e N_D} \left(U - U_{FB} - \frac{kT}{e} \right) \quad (10)$$

where C is the measured interfacial capacitance, C_H is the capacitance of the Helmholtz layer, ε_r is the relative permittivity or dielectric constant of the WO_3 nanostructures (a value of 50 has been assumed for the WO_3 [9,22,35,36]), ε_0 is the vacuum permittivity ($8.85 \times 10^{-14} \text{ F/cm}$), e is the electron charge ($1.60 \times 10^{-19} \text{ C}$), U is the applied potential, k is the Boltzmann constant ($1.38 \times 10^{-23} \text{ J/K}$) and T is the absolute temperature.

Values of N_D and U_{FB} can be determined from the slope and the intercept, respectively, of a linear representation between C^{-2} and U , according to Eq. (10). It is noteworthy that C_H does not influence the value of N_D , but it affects the value of U_{FB} . It is usually assumed that the Helmholtz layer capacitance is so large compared with the space charge layer capacitance that the total capacitance measured can be treated as the space charge layer capacitance and the potential drop caused by the applied potential occurs entirely within the space charge region. However, in highly doped semiconductors (such as nanostructured metal oxides), C_H should not be neglected, since a significant part of the potential difference at the semiconductor/electrolyte interface extends to the Helmholtz layer in the solution [46–48]. Therefore, in the present study, C_H values have not been neglected. For practical purposes, since the electrolyte used in electrochemical measurements (1 M HCl) was sufficiently concentrated, the capacitance response of the Guoy-Chapman diffuse layer can be ignored, so $C_H \approx C_{\text{nano}}$.

Table 1 shows the values of N_D and U_{FB} for the WO_3 nanostructures synthesized at different rotation velocities. As mentioned above, U_{FB} values have been determined by taking into account the capacitance of the Helmholtz layer in Eq. (10). It can be seen that N_D were very high, of the same order of magnitude than values reported in the literature for WO_3 nanoplatelets, nanoparticles and nanorods [9,22,35,36,49]. Samples anodized at 250, 375 and 500 rpm presented significantly higher donor densities than the rest of the samples, which can be explained by an increase of the interfacial capacitance due to the larger surface area exposed to the electrolyte for the samples anodized at intermediate rotation velocities, resulting in higher concentration of oxygen vacancies. These results are consistent with the increase in C_{nano} and the decrease in R_{nano} (increase in the nanostructures conductivity) observed above for these samples (Fig. 6). Concerning the U_{FB} , there was no clear

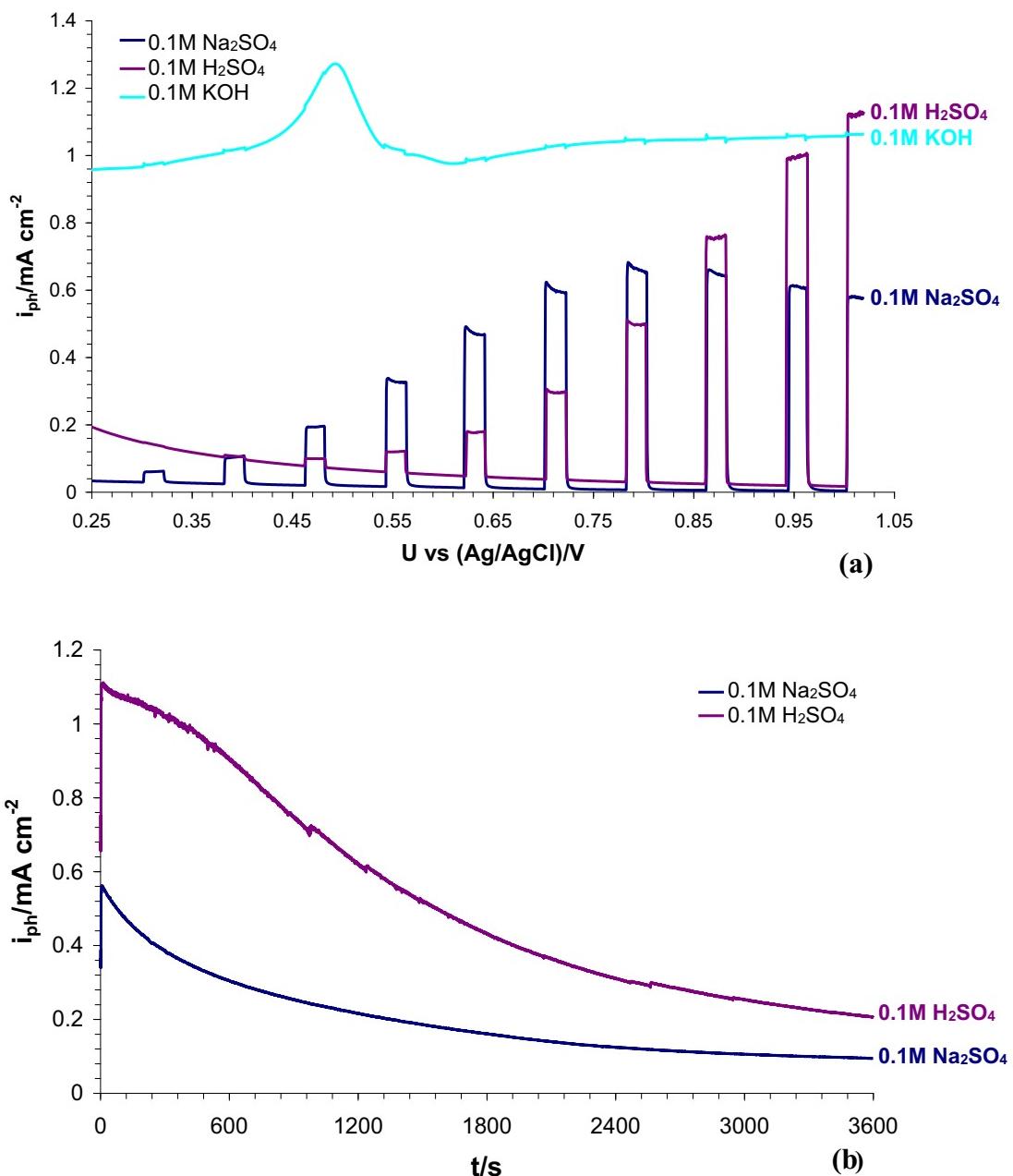


Fig. 8. (a) Photocurrent density (i_{ph}) vs. potential (U) curves of the WO_3 nanostructure formed at 500 rpm measured in 0.1 M Na_2SO_4 , 0.1 M H_2SO_4 and 0.1 M KOH solutions under AM1.5 illumination. (b) Photocurrent density transients under illumination obtained at an applied potential of 1.02 $\text{V}_{\text{Ag}/\text{AgCl}}$ in the 0.1 M Na_2SO_4 and 0.1 M H_2SO_4 electrolytes.

tendency with rotation velocity, and values varied within a small range from 0.23 to 0.30 $\text{V}_{\text{Ag}/\text{AgCl}}$, similar to values obtained by other authors [22,35,36].

3.4. Degradation resistance of the samples

As it has been mentioned in the introduction section, resistance to degradation in a given environment under illumination is a critical requirement for the photoanode material. The first step is to ensure that the semiconductor oxide used as photoanode does not undergo chemical and electrochemical degradation in the dark. In general, it is not recommended to work with WO_3 at pH higher than 4–5, since it is thermodynamically unstable and dissolves [33,34]. Fig. 8a shows the photocurrent density-potential (i_{ph} - U) curves for the sample anodized at 500 rpm, recorded in different electrolytes

with different pH values. The sample synthesized at 500 rpm has been used throughout this section by way of illustration, and the conclusions obtained here can be perfectly extrapolated to the rest of the samples. It can be observed that the shape of the photocurrent density response of samples tested in 0.1 M Na_2SO_4 and 0.1 M H_2SO_4 solutions was similar, although in the latter case, photocurrent density values were higher at potentials $U > 0.8 \text{ V}$, while dark current densities were stable and very low (of the order of few $\mu\text{A cm}^{-2}$ or even lower). These results denote that WO_3 nanostructures were resistant to chemical and electrochemical degradation under dark conditions in the 0.1 M Na_2SO_4 and 0.1 M H_2SO_4 electrolytes, especially in the acidic one.

On the other hand, the photocurrent density-potential curve recorded in 0.1 M KOH is completely different. Although the curve

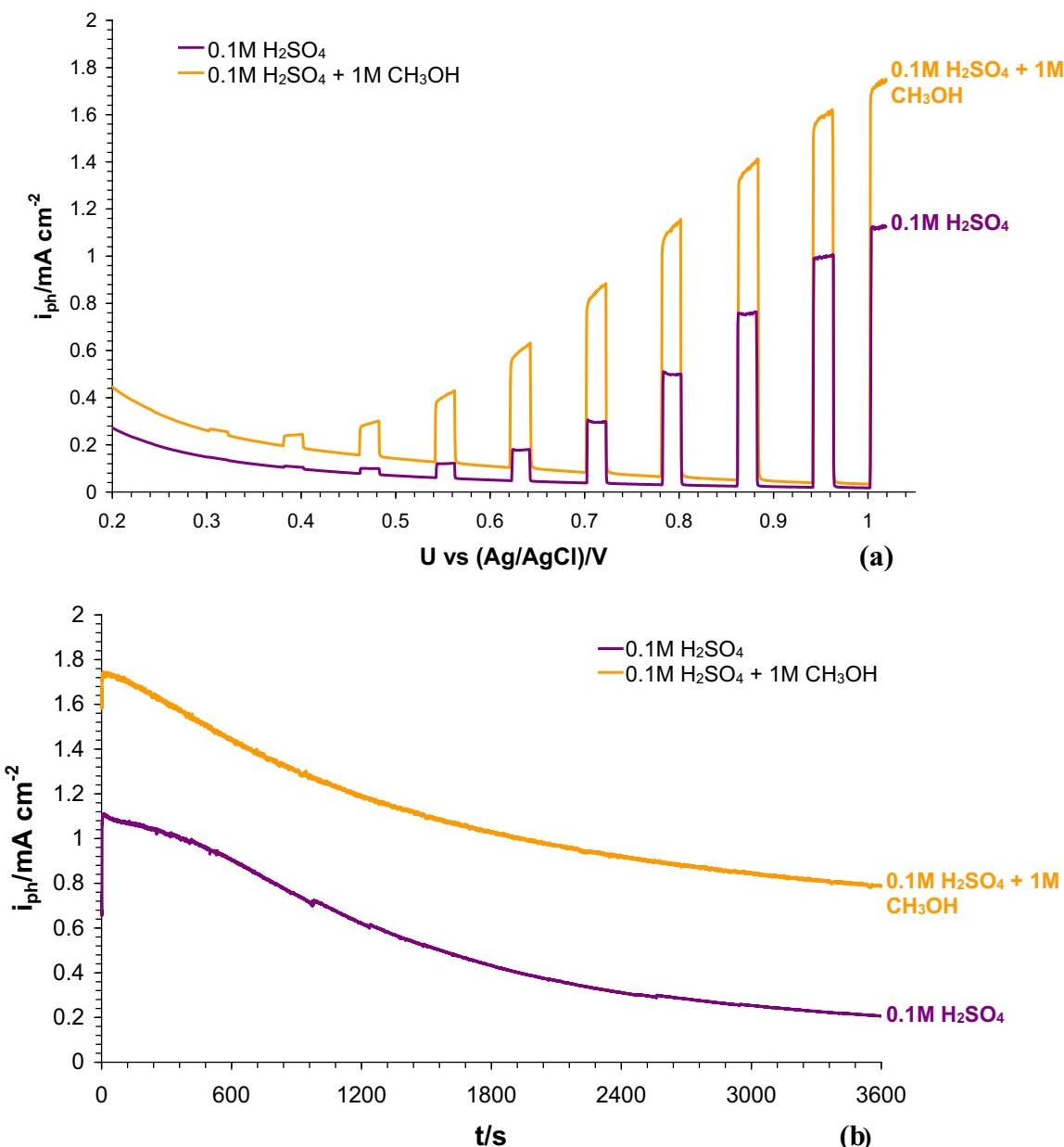
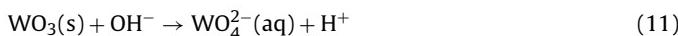


Fig. 9. (a) Photocurrent density (i_{ph}) vs. potential (U) curves of the WO_3 nanostructure formed at 500 rpm measured in $0.1\text{ M H}_2\text{SO}_4$ and $0.1\text{ M H}_2\text{SO}_4 + 1\text{ M CH}_3\text{OH}$ solutions under AM1.5 illumination. (b) Photocurrent density transients under illumination obtained at an applied potential of $1.02\text{ V}_{\text{Ag}/\text{AgCl}}$ in the same electrolytes.

was also obtained under chopped illumination, it was difficult to distinguish between photocurrent and dark current densities (Fig. 8a). The appearance of a wide peak centered at $\sim 0.5\text{ V}$ indicates that some kind of transformation in the characteristics or composition of the WO_3 surface film took place. Moreover, after the photoelectrochemical test, the WO_3 nanostructure was completely dissolved from the surface in contact with the KOH electrolyte. These results imply that WO_3 nanoplatelets were unstable in alkaline media, as expected, due to the OH^- -induced chemical dissolution, in accordance with [32,34,49]:



According to Anik and Cansizoglu [32], the chemical dissolution of the oxide film given in Eq. (11) is the rate-determining step in the oxidation of W to WO_3 . This fact can explain the very high current densities obtained in the 0.1 M KOH electrolyte, which correspond to the active oxidation of the W substrate.

Concerning the photostability of the WO_3 nanostructures synthesized at 500 rpm, Fig. 8b shows the photocurrent density transients obtained at an applied potential of 1.02 V in the $0.1\text{ M Na}_2\text{SO}_4$ and $0.1\text{ M H}_2\text{SO}_4$ electrolytes. It can be observed that photocurrent density values considerably decreased after 1 h of exposition to simulated sunlight, regardless of the electrolyte: a decrease of 83.9% (from 0.56 to 0.09 mA cm^{-2}) was observed in the case of the $0.1\text{ M Na}_2\text{SO}_4$ solution, and a decrease of 81.1% (from 1.11 to 0.21 mA cm^{-2}) the $0.1\text{ M H}_2\text{SO}_4$ electrolyte. Other authors observed a similar drop in photocurrent densities with time working with WO_3 photoanodes in acidic or neutral electrolytes [30,35–38,49,50].

Although it has been demonstrated that WO_3 nanoplatelets were chemical and electrochemically stable in the acidic and neuter solutions, the fast decrease in photocurrent densities with time in these electrolytes indicates that samples underwent degradation under illumination. Photocorrosion is a phenomenon that takes

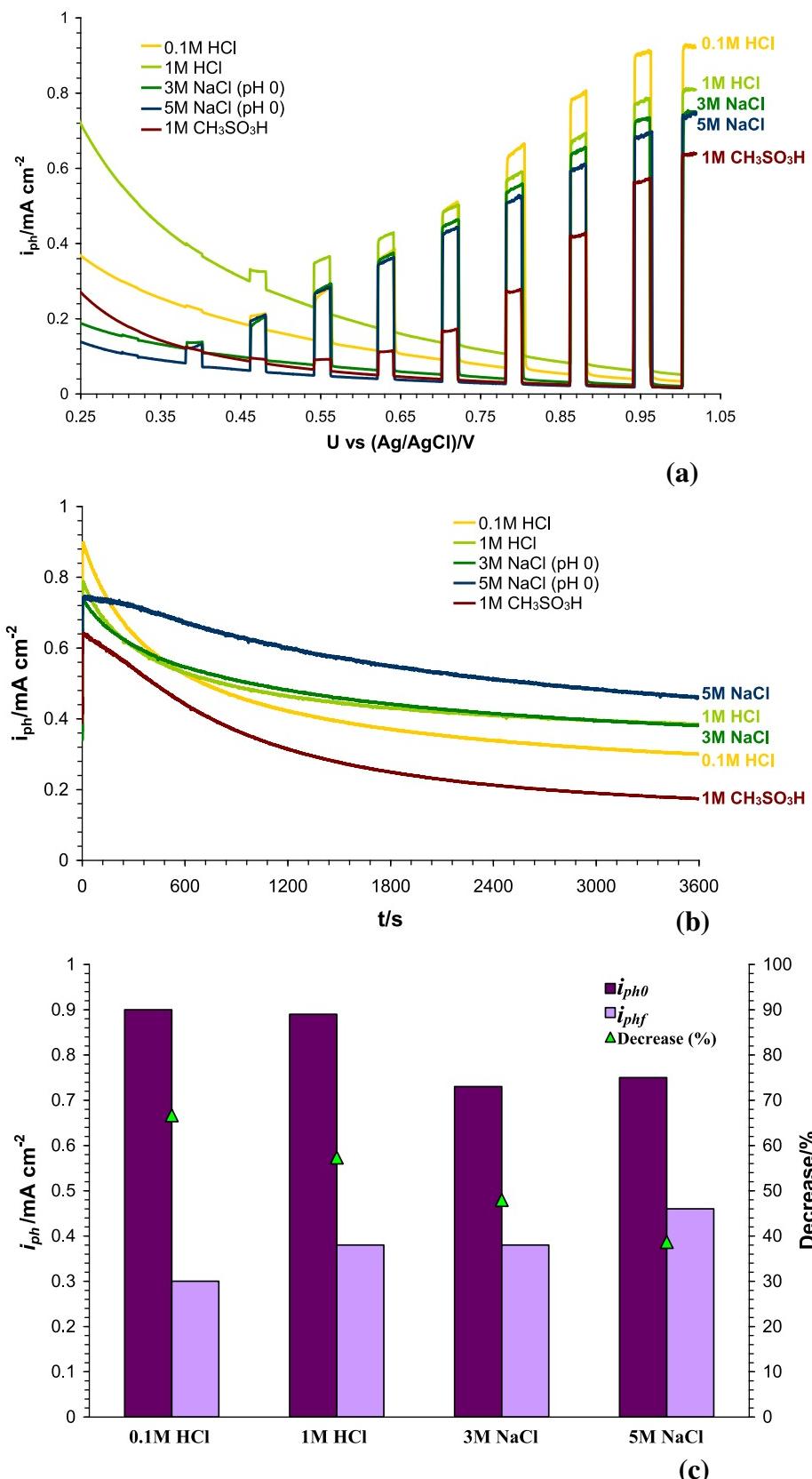


Fig. 10. (a) Photocurrent density (i_{ph}) vs. potential (U) curves of the WO_3 nanostructure formed at 500 rpm measured in 0.1 M HCl, 1 M HCl, 3 M NaCl, 5 M NaCl and 1 M $\text{CH}_3\text{SO}_3\text{H}$ solutions under AM1.5 illumination. (b) Photocurrent density transients under illumination obtained at an applied potential of 1.02 $\text{V}_{\text{Ag}/\text{AgCl}}$ in the same electrolytes. (c) Values of photocurrent density before and after the stability tests and % of photocurrent density decrease in the same electrolytes as before.

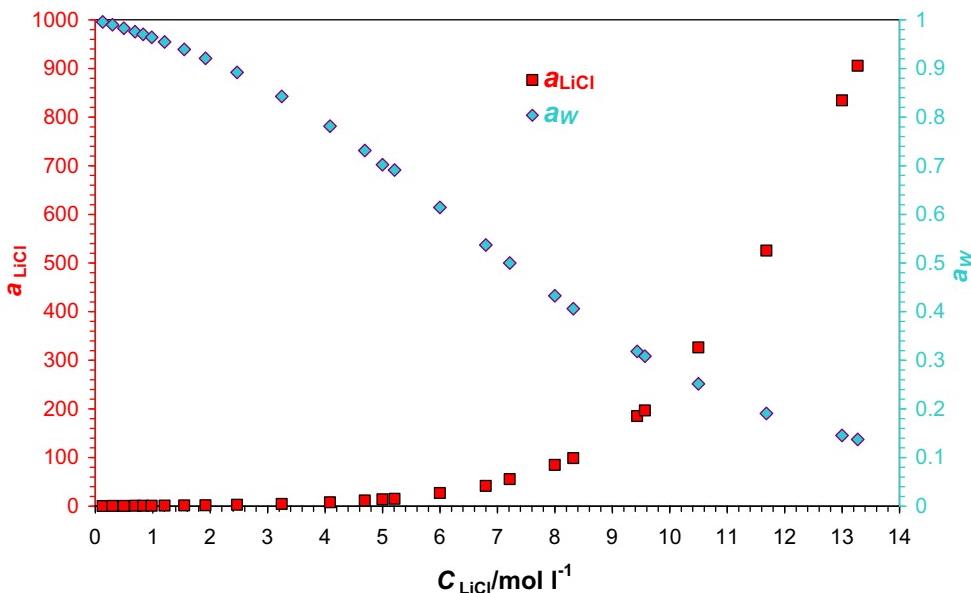
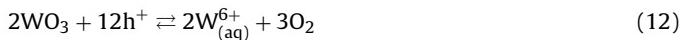
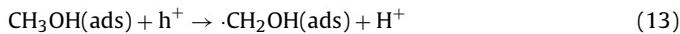


Fig. 11. Variation of activity of chloride ions (a_{LiCl}) and activity of water (a_w) with the LiCl molar concentration.

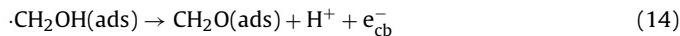
place when the photogenerated holes in the valence band oxidize the semiconductor itself, instead of species present in the electrolyte [16,51]. In the case of WO_3 , photocorrosion reaction proceeds through the following reaction:



Moreover, if holes accumulate near the semiconductor surface instead of reacting with species in the electrolyte, photocurrent density can also decrease with time due to the recombination of part of the photogenerated charge carriers (electrons and holes) [52]. A solution for these two issues (photocorrosion and recombination) can be the addition of a hole scavenger (electron donor) to the electrolyte, such as methanol [8,9,53]. The hole scavenger reacts with the holes at the semiconductor/electrolyte interface, thus reducing the complications arising from hole accumulation. Fig. 9a shows the photocurrent density-potential curve for the sample anodized at 500 rpm, recorded in a 0.1 M H_2SO_4 electrolyte in the absence and presence of 1 M methanol. An increase in photocurrent density values can be clearly observed in the 0.1 M $\text{H}_2\text{SO}_4 + 1 \text{ M CH}_3\text{OH}$ solution, which indicates hole scavenging by CH_3OH [8,9,53], according to the following reaction [53,54]:

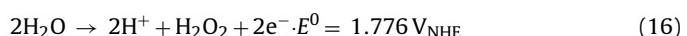
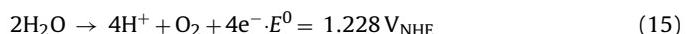


Besides, the adsorbed intermediate radical $\cdot\text{CH}_2\text{OH}(\text{ads})$, generated during the oxidation of methanol, can also effectively inject an additional electron into the conduction band of WO_3 :



which explains the important increase observed in photocurrent density in the presence of CH_3OH [8,9,53]. However, these high photocurrent density values recorded in the presence of CH_3OH tended to gradually decrease with increasing irradiation time, as in the case of the electrolyte without methanol (Fig. 9b). In this case, the photocurrent density decrease was of 54.6% (from 1.74 to 0.79 mA cm^{-2}) in one hour. Although this photocurrent density drop is lower than that recorded in the 0.1 M H_2SO_4 solution, this result demonstrates that neither photocorrosion due to hole accumulation at the interface nor recombination of photogenerated electron-hole pairs were the main cause of the fast photocurrent density decrease with time under simulated sunlight.

The rapid deactivation of the WO_3 electrodes upon illumination must be ascribed to a different light-driven degradation mechanism. Several authors have reported the occurrence of photocorrosion of WO_3 induced by the formation of H_2O_2 during incomplete water oxidation (a reaction that is kinetically favored over oxygen evolution reaction (OER) in WO_3 photoanodes under illumination, in spite of being thermodynamically less feasible [34,37,55–58]):



Subsequently, hydrogen peroxide interacts with the WO_3 nanostructure, and peroxy-species accumulates on the surface, partially blocking it and thus reducing photoactivity [26,37,38,53]. These tungsten peroxide species can also induce the recombination of photogenerated electron-hole pairs at the surface of the WO_3 nanostructures [38]. Hydroxyl radicals may form during the process, as well [57]:



which can subsequently interact with the WO_3 surface [36,57], leading to the same negative consequences for the photoanode stability.

On the other hand, the oxidation of some acid counterions (Cl^- , SO_4^{2-} , CH_3SO_3^- , etc.) can also take place on the photoanode surface under illumination, since holes in the valence band of WO_3 have a potential of $\sim 2.97 \text{ V}_{\text{NHE}}$ at pH 0 [56,57,59], enough to drive these oxidation reactions ($E^0(\text{Cl}_2/\text{Cl}^-) = 1.358 \text{ V}_{\text{NHE}}$; $E^0(\text{S}_2\text{O}_8^{2-}/\text{SO}_4^{2-}) = 2.010 \text{ V}_{\text{NHE}}$; $E^0((\text{CH}_3\text{SO}_3)_2/\text{CH}_3\text{SO}_3^-) = 2.12 \text{ V}_{\text{NHE}}$ [60]). Hence, the oxidation of acid counterions can kinetically compete with peroxide formation and OER [34,56,57,60].

Hill and Choi [56] observed that the oxidation of chloride anions was the dominant photooxidation reaction taking place on the WO_3 surface when immersing the photoanode in a 0.1 M NaCl solution (pH 1–5) under illumination. They also observed that photocurrent density values did not significantly decrease over time, which they ascribed to the effective suppression of peroxide formation by chloride oxidation. On the other hand, Augustynski and co-workers recently reported a very good photoelectrochemical

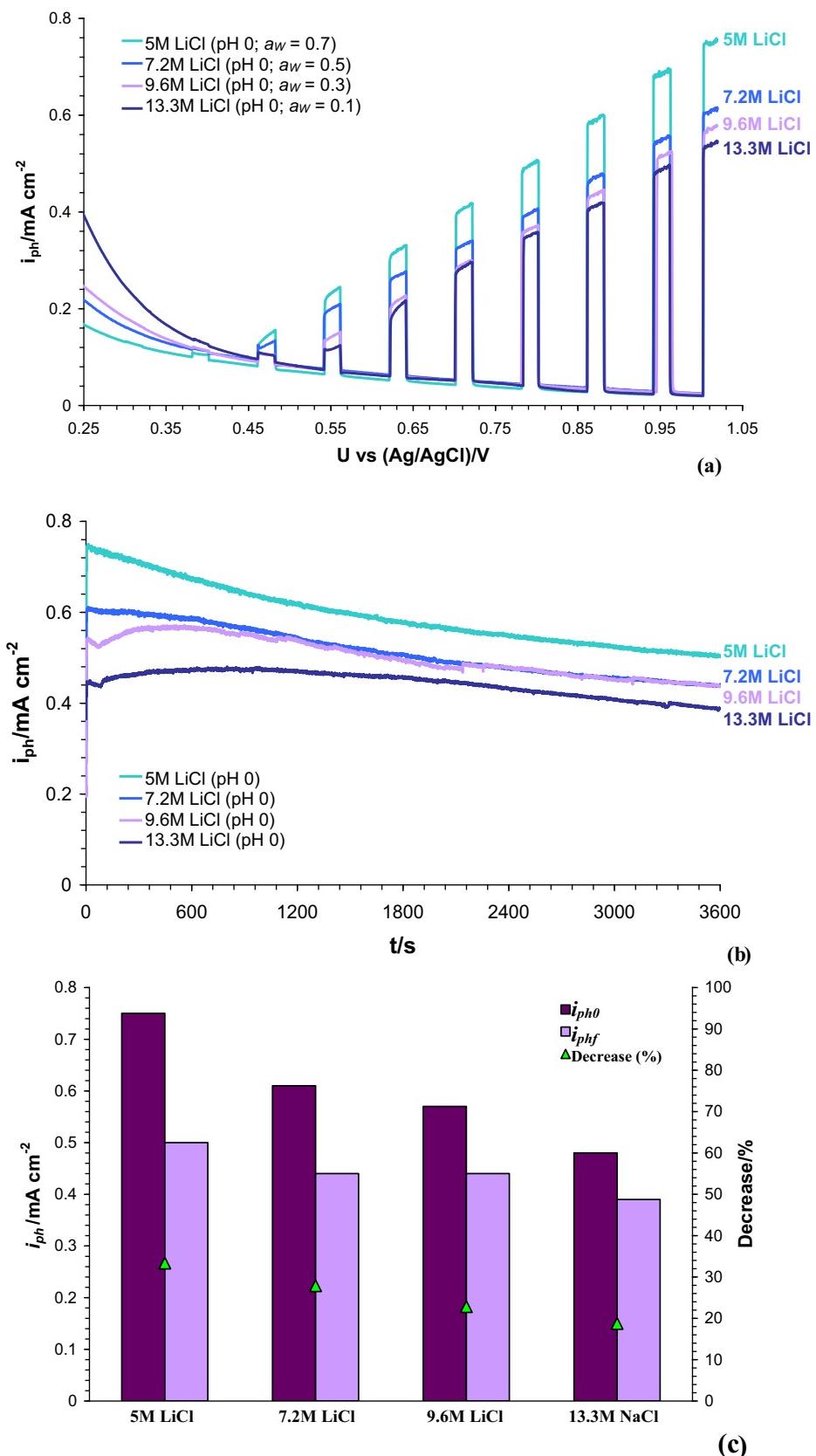


Fig. 12. (a) Photocurrent density (i_{ph}) vs. potential (U) curves of the WO_3 nanostructure formed at 500 rpm measured in 5 M LiCl, 7.2 M LiCl, 9.6 M LiCl and 13.3 M LiCl solutions under AM1.5 illumination. (b) Photocurrent density transients under illumination obtained at an applied potential of $1.02 \text{ V}_{\text{Ag}/\text{AgCl}}$ in the same electrolytes. (c) Values of photocurrent density before and after the stability tests and % of photocurrent density decrease in the same electrolytes as before.

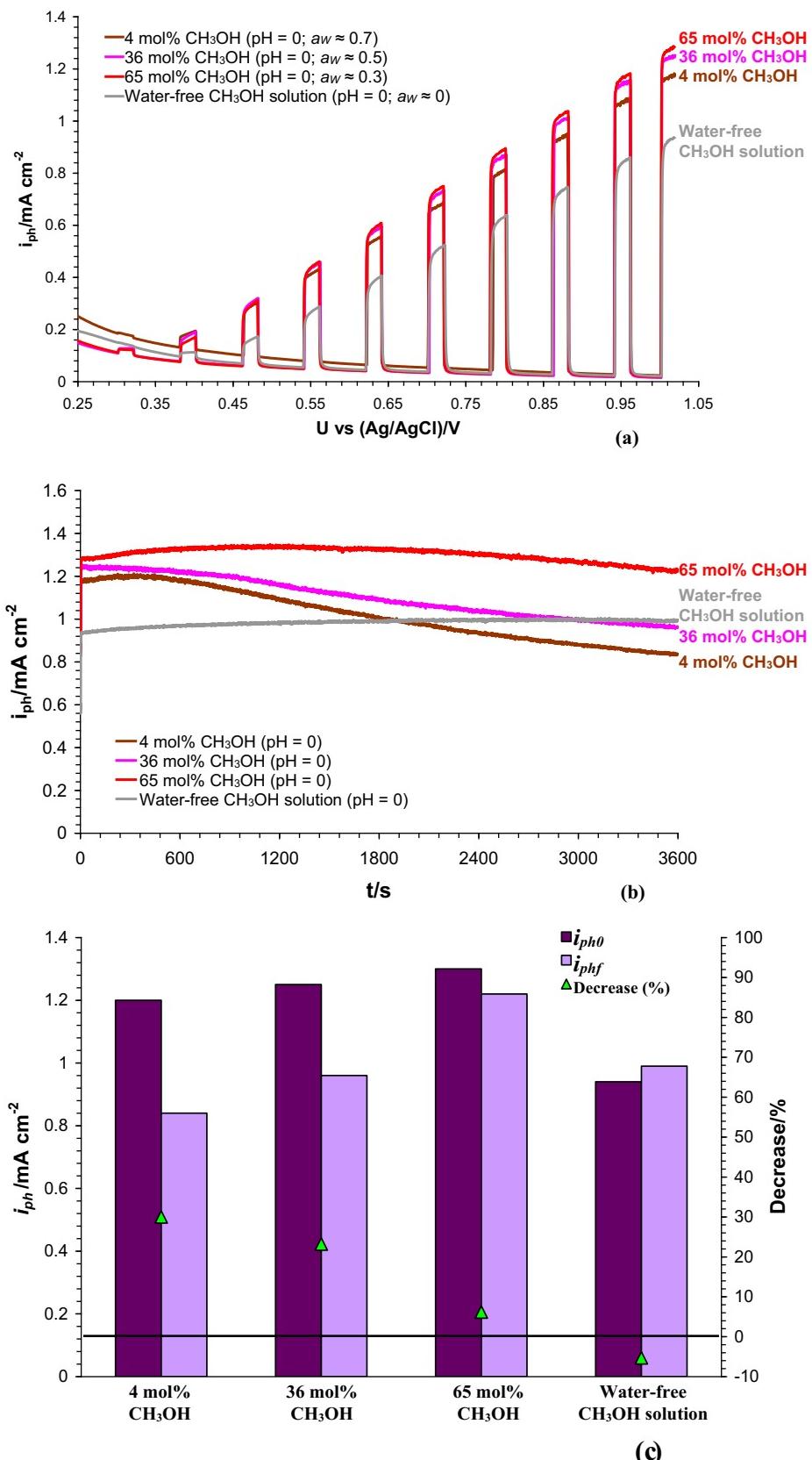


Fig. 13. (a) Photocurrent density (i_{ph}) vs. potential (U) curves of the WO_3 nanostructure formed at 500 rpm measured in different aqueous/methanolic solutions with 5 M LiCl under AM1.5 illumination. (b) Photocurrent density transients under illumination obtained at an applied potential of $1.02 \text{ V}_{\text{Ag}/\text{AgCl}}$ in the same electrolytes. (c) Values of photocurrent density before and after the stability tests and % of photocurrent density decrease in the same electrolytes as before.

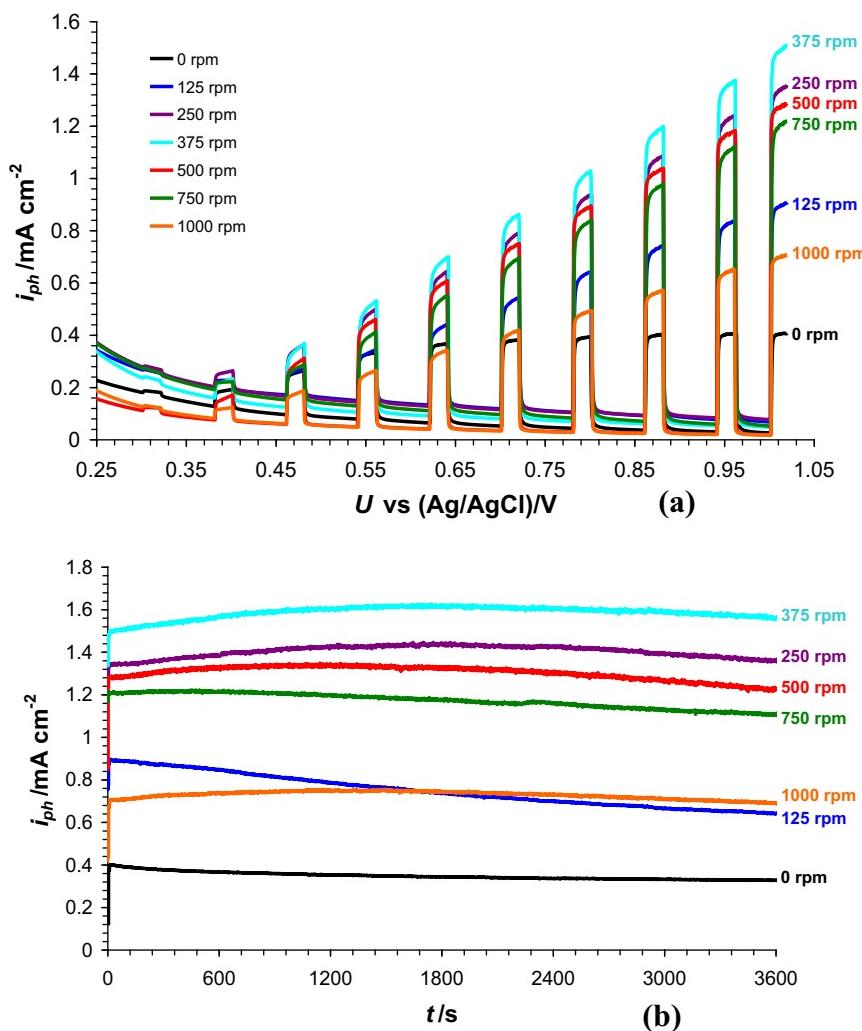


Fig. 14. (a) Photocurrent density (i_{ph}) vs. potential (U) curves of the WO_3 nanoplatelets formed at different rotation velocities, measured in a 65 mol% CH_3OH with 5 M LiCl solution under AM1.5 illumination. (b) Photocurrent density transients under illumination obtained at an applied potential of $1.02 \text{ V}_{\text{Ag}/\text{AgCl}}$ in the same electrolyte.

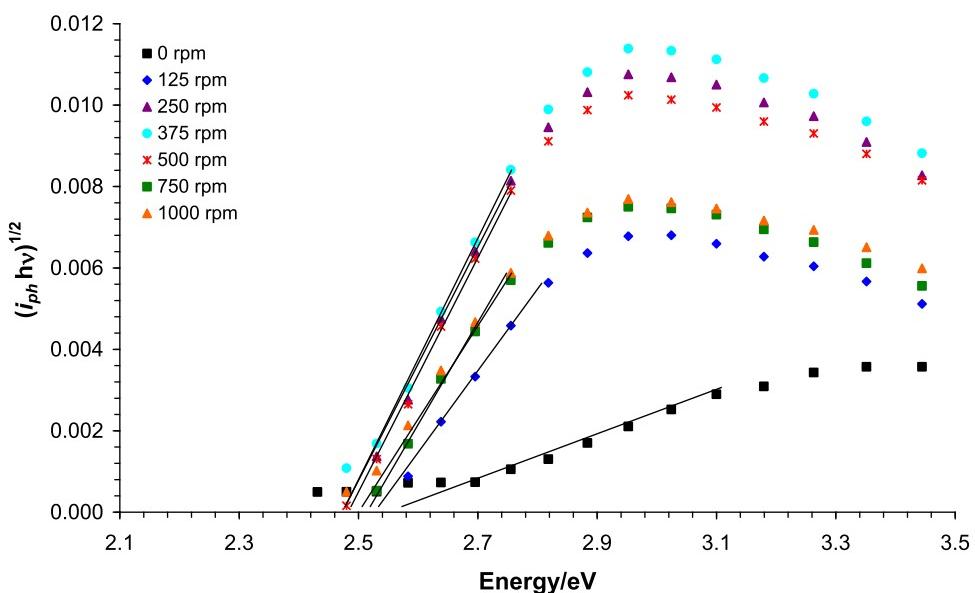


Fig. 15. Band-gap determination for the WO_3 nanoplatelets formed at different rotation velocities from the $(i_{ph} h\nu)^{0.5}$ vs. $h\nu$ plots.

behavior of nanostructured WO_3 photoanodes in a methanesulfonic acid ($\text{CH}_3\text{SO}_3\text{H}$) electrolyte [38], achieving large and stable photocurrent densities. They ascribed this good behavior to the minimization of the $e^- - h^+$ recombination coming from the formation of a surface layer of tungsten peroxy-species, although the mechanism by which this was achieved in $\text{CH}_3\text{SO}_3\text{H}$ solutions in particular was not clearly explained.

In order to investigate a way of suppressing or, at least, lessening the negative effects of the photo-induced generation of hydrogen peroxide on the stability of the WO_3 photoanodes, i_{ph} - U curves and stability tests were also performed in HCl, NaCl (pH adjusted to 0 by adding small amounts of highly concentrated (37%) HCl) and $\text{CH}_3\text{SO}_3\text{H}$ electrolytes (Fig. 10). In the case of the sample immersed in the methanesulfonic acid solution, it can be observed that recorded photocurrent densities were lower than those obtained in the electrolytes with Cl^- , regardless of the chloride concentration and the applied potential (Fig. 10a). Moreover, it is evident from Fig. 10b that WO_3 nanostructures were not stable in that solution under illumination (the photocurrent density decreased a 73% after 1 h, a value similar to those reported above for 0.1 M Na_2SO_4 and 0.1 M H_2SO_4 electrolytes).

On the other hand, for the Cl^- containing electrolytes, it can be observed that photocurrent densities slightly decreased at 1.02 V as the chloride concentration increased (Fig. 10a). This fact can be related to some interaction between the WO_3 nanostructure and the Cl^- anions, which led to a (minor) reduction in photoactivity. Regarding the stability of the photoanodes, WO_3 samples were more stable in concentrated chloride solutions (i.e. 5 M NaCl) than in diluted ones (Fig. 10b), since the decrease experienced by photocurrent density in the 5 M NaCl solution was the lowest (39%) (Fig. 10c). This fact can be explained by taking into account two factors. As it has been explained above, Cl^- anions competed with H_2O for the valence band holes; thus, the oxidation of Cl^- to Cl_2 somewhat limited the incomplete water oxidation to hydrogen peroxide and hence moderately improved the photostability of the samples [56]. However, little differences can be discerned in the photocurrent density response of the samples immersed in the 1 M HCl and 3 M NaCl (pH 0), but this response was different in the 5 M NaCl (pH 0) electrolyte, where the decrease in photocurrent density with time was not abrupt (Fig. 10b). Therefore, the presence of Cl^- anions competing with water for the valence band holes was not the only factor that positively influenced the photoanode stability. The other factor was the reduction in the activity of water (a_w). In diluted or moderately concentrated electrolytes, a_w usually takes values very close to one; in highly concentrated solutions, however, a_w falls below unity, since hydration of ions removes water from the “role” of solvent to become a part of the hydrated ion, thus reducing its thermodynamic activity. Hence, a reduction in a_w with increasing electrolyte concentration can be understood as a reduction in the amount of free H_2O molecules in the solution. Since water was the reacting species in the formation of hydrogen peroxide and hydroxyl radicals, according to Eqs. (16) and (17), a reduction in a_w led to a reduction in the formation of H_2O_2 and $\cdot\text{OH}$.

From osmotic coefficient values (φ) of NaCl solutions [61,62], a_w has been calculated according to the next equation [63,64]:

$$\ln(a_w) = -\phi M_w \zeta m \quad (18)$$

where M_w is the molar mass of water and ζ is the stoichiometric parameter, i.e., the number of ion moles produced by complete dissociation of one mole of salt (for a 1:1 salt such as NaCl or LiCl, ζ equals 2); notice that molality m is expressed in kilomole of solute per kilogram of solvent (or mole of solute per gram of solvent) so the units in Eq. (18) are consistent. In the present case, a_w for the 3 M NaCl electrolyte is 0.88, while a_w for the 5 M NaCl electrolyte is 0.78. Consequently, this decrease in a_w contributed to stabilize to

some extent the photoanode immersed in the latter electrolyte by reducing the formation of hydrogen peroxide and hydroxyl radicals through Eq. (16) and Eq. (17).

Since a decrease in a_w has been observed to positively affect the photoresponse of the WO_3 nanostructures in terms of an improvement of their photodegradation resistance, additional photoelectrochemical tests were carried out in highly concentrated chloride solutions. In NaCl electrolytes, a_w cannot be decreased below a value of ~0.75 [61], so LiCl solutions were used instead for that purpose. Given that the Li^+ cation is smaller than the Na^+ one but both have the same charge, the smallest cation produces the strongest electrical field, resulting in stronger interactions with the surrounding water molecules. Thus, working with LiCl solutions, a_w can be further reduced [65]. In Fig. 11, the variation of a_w with molar concentration for LiCl electrolytes is plotted. It can be observed that a_w continuously decreases as LiCl concentration increases, this effect being more marked at high concentrations, where a_w values are very low, close to zero. In this figure, activity of Cl^- anions is also represented against the molar concentration of LiCl. This activity, a_{LiCl} , has been calculated from the following equation:

$$a_{\text{LiCl}} = \frac{m_{\text{LiCl}}}{m^0} \gamma_{\pm} \quad (19)$$

where the LiCl molality, m_{LiCl} , is defined as mol LiCl/kg H_2O and m^0 is the standard LiCl molality ($m^0 = 1$ mol LiCl/kg H_2O). Activity coefficients and osmotic coefficients have been obtained from the literature [61], as well as density values of LiCl aqueous electrolytes used to transform molality to molarity [66].

Fig. 12a shows the i_{ph} - U curves of the sample anodized at 500 rpm and immersed in several concentrated LiCl electrolytes. It can be observed that i_{ph} values decreased to some extent as the LiCl concentration increased, as it was also observed in Fig. 10a. This fact may be explained by some negative interaction between chlorides and WO_3 , taking into account that activity of Cl^- anions sharply increased for LiCl concentrations higher than 5 M (Fig. 11). On the other hand, in spite of the lower photocurrent density values, it is evident that WO_3 nanostructures became more stable as LiCl concentration increased (Figs. 12b and c), due to the reduction of a_w , as explained above: in the 5 M LiCl electrolyte, $a_w = 0.7$, while in the 13.3 M LiCl solution $a_w = 0.1$, so the generation of H_2O_2 and hydroxyl radicals via Eqs. (16) and (17), respectively, was significantly reduced in the most concentrated LiCl electrolyte, this reduction positively affecting the stability of the samples.

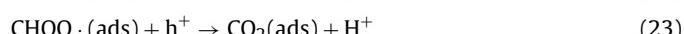
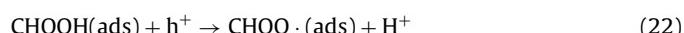
Nonetheless, and despite the better performance of WO_3 nanostructures in low a_w electrolytes in terms of its photodegradation resistance, photocurrent density values obtained from the i_{ph} - U curves were appreciably lower in highly concentrated Cl^- solutions than, for example, in 0.1 M H_2SO_4 . In consequence, the next step was to keep the values of a_w low, thus ensuring the good stability of the samples, but to increase at the same time photocurrent densities. The ability of alcohols to reduce the activity of water is well known [67,68]. On the other hand, a high ionic conductivity within the WO_3 nanostructures is crucial to prevent uneven current and potential distributions across the nanostructured semiconductor electrode [38], which would negatively affect its photoelectrochemical performance. Methanol possesses an electrical conductivity higher than other alcohols [69] and, as it has been observed above (Fig. 9), the role of methanol as a hole scavenger resulted in a considerable increase in photocurrent density due to the direct injection of an electron to the WO_3 conduction band (Eq. (14)). In that sense, a_w has been reduced by adding different amounts of CH_3OH to a 5 M LiCl aqueous electrolyte: 4 mol% CH_3OH (1 M; $a_w \approx 0.7$), 36 mol% CH_3OH ($a_w \approx 0.5$), 65 mol% CH_3OH ($a_w \approx 0.3$), water-free CH_3OH solution ($a_w \approx 0$). Activity of water values in these solutions have been estimated

from the literature [67,68], taking the values of a_W shown in Fig. 11 for aqueous LiCl solutions as starting point. In the i_{ph} -U curves for the different CH₃OH-LiCl electrolytes (Fig. 13a) it can be observed that photocurrent densities increased as the mol% of CH₃OH increased, except for the water-free CH₃OH solution, in which i_{ph} took the lowest values of the series. Moreover, photocurrent densities were considerably higher than those obtained in LiCl aqueous solutions (Fig. 12a), which can be explained by the additional electron injected in the WO₃ valence band during the oxidation of the adsorbed intermediate radical, •CH₂OH(ads) (Eq. (14)). In the water-free CH₃OH solution, the decrease observed in i_{ph} was most likely due to a decrease in the ionic conductivity of the solution (the conductivity of water-free 5 M LiCl/methanol is 11.89 mS cm⁻¹, while that of 5 M LiCl/65 mol% methanol/water is 22.6 mS cm⁻¹).

Concerning the stability of the WO₃ nanostructures immersed in these electrolytes, an increase in the mol% of CH₃OH increased the resistance of the samples to photodegradation, since the decrease (in%) of i_{ph} in 1 h was manifestly reduced to a value below 10% in the 65 mol% CH₃OH solution; in the water-free CH₃OH solution i_{ph} even increased in 5.3%, although photocurrent densities were lower than in the 65 mol% CH₃OH solution. These results confirm the beneficial influence that a reduction in a_W had on the resistance of the WO₃ nanostructures to undergo H₂O₂-driven photodegradation. Moreover, the reduction in a_W achieved by adding CH₃OH to the 5 M LiCl solution avoided the negative effect that elevated chloride activities had on the photoelectrochemical performance of the samples, without leading to an excessive decrease in ionic conductivity, due to the high concentration of LiCl.

3.5. Photoelectrochemical activity

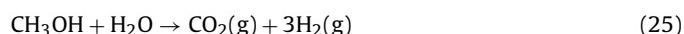
Photocatalytic H₂ production from methanol-containing solutions has previously been reported in the literature [54,70–73]. In these cases, H₂ was the main gaseous product evolving on the cathode, while CO₂ evolved on the anode after more or less prolonged irradiation due to the complete oxidation of CH₃OH [54,71]. The direct oxidation of methanol by surface holes starts with Eqs. (13) and (14) and continues as follows [54]:



The complete oxidation of methanol via Eqs. (13), (14), (20)–(23) involves the formation of 6H⁺ per mol of CH₃OH, so the cathodic reaction is:



And the overall reaction is:



These overall methanol decomposition has a lower splitting energy than water (0.7 eV vs. 1.23 eV) [72]. In a recent work by Lin et. al [72], it was observed that the maximum hydrogen production using a Pt/TiO_{2-x}N_x catalyst was achieved in a solution 80 vol% of methanol, which corresponds approximately to a 65 mol% CH₃OH. Hence, in concentrated methanolic aqueous solutions, the photo-generation of H₂ seems to be enhanced, while CO₂ evolves on the anode. In the present case, where the methanol/water solution also has 5 M LiCl, gaseous Cl₂, as well as some O₂, may also form on the anode.

Fig. 14 shows the i_{ph} -U curves and the photocurrent density transients recorded for 1 h under simulated sunlight of the different

Table 2
Band-gap values of the WO₃ nanoplatelets formed at different rotation velocities.

Sample (rpm)	E_g /eV
0	2.55 ± 0.04
125	2.53 ± 0.03
250	2.49 ± 0.02
375	2.48 ± 0.03
500	2.49 ± 0.04
750	2.53 ± 0.05
1000	2.52 ± 0.06

WO₃ samples anodized at various rotation velocities, immersed in the 5 M LiCl with 65 mol% CH₃OH electrolyte. It can be observed that the relationship between photocurrent density values and rotation velocity is not direct. The sample anodized at 375 rpm provided the highest photoresponse, which is consistent with this sample having the largest electrochemically active surface area. This sample was followed by those synthesized at 250, 500 and 750 rpm (in that order), which also presented high photocurrent densities, while the photoresponse of the samples anodized at 125 and 1000 rpm was significantly lower. Finally, the WO₃ nanoplatelets formed under static flow conditions (0 rpm) provided the lowest photocurrent density. These results are consistent with the previous findings, since samples synthesized at low (0 and 125 rpm) and high (1000 rpm) rotation velocities presented lower electrochemically active surface areas and higher charge-transfer resistance values than those anodized at intermediate rpm (250–500 rpm).

Fig. 15 shows the $(i_{ph} h\nu)^{0.5}$ vs. $h\nu$ plots for the different samples obtained by measuring the photocurrent density of the WO₃ nanoplatelets at an applied potential of 1.02 V_{Ag/AgCl} in the wavelength range from 300 to 500 nm. These plots have been used to determine the indirect band-gap energy (E_g) of the nanostructures, where h is Planck's constant (4.1357×10^{-15} eV s) and ν is the frequency of light. The E_g values, obtained from the x-intercept of the extrapolated linear part of the diagram, are shown in Table 2. The calculated E_g are slightly lower than the typical value for crystalline WO₃ (2.6 eV) [6]. This slight red shift can be explained by an enhancement in light scattering due to the nanoplatelet morphology resulting in higher light adsorption by the nanostructures [30]. No appreciable differences in E_g were observed for samples anodized at different rotation velocities.

4. Conclusions

Globular clusters of WO₃ nanoplatelets associated in a tree-like fashion have been obtained by anodizing W in an acidic F-containing electrolyte under controlled hydrodynamic conditions. The thickest nanoplatelets layers were obtained at 375 rpm, since this rotation velocity was high enough to enhance mass transfer towards the electrode surface and hence to increase the formation of soluble species and their subsequent precipitation in the form of hydrated WO₃, but it was not too high as to facilitate the removal of these soluble species from the electrode surrounding area.

The charge-transfer resistance at the nanostructure/electrolyte interface decreased for the samples anodized at intermediate rotation velocities (250–500 rpm), while the double layer capacitance increased for these samples, which can be explained by an increase of their electrochemically active surface area.

Mott-Schottky analysis concluded that tree-like WO₃ nanoplatelets synthesized at intermediate rotation velocities (250–500 rpm) had the highest concentration of oxygen vacancies in their crystalline structures, resulting in higher electrical conductivities, which is consistent with the decrease in the charge-transfer resistance observed through EIS measurements.

WO_3 nanostructures were not stable in alkaline media due to the active dissolution of the oxide by the action of OH^- . Moreover, these nanostructures were not resistant to photodegradation in aqueous acidic electrolytes with low ionic strengths due to the photogeneration of peroxides and hydroxyl radicals that interacted with the surface of the electrode and resulted in a gradual loss of photoactivity. In highly concentrated chloride solutions the stability of the samples manifestly improved due to the reduction of the activity of water, which led to a decrease in the formation of peroxides and hydroxyl radicals. However, the photoactivity of the samples decreased with increasing chloride activity, so aqueous/methanolic mixed solutions with high chloride concentration were used in the end to assure a very good photostability of the samples.

Apart from being a hole scavenger and helping in the separation of photogenerated electron-hole pairs, methanol has been demonstrated to play an important role in reducing the activity of water and thus the H_2O_2 -induced photodegradation of WO_3 nanoplatelets. The suppression of peroxide species remarkably improved the long-term photostability of WO_3 .

Photocurrent densities measured indicated the superior photocatalytic activity of tree-like WO_3 nanoplatelets synthesized at intermediate rotation velocities, especially at 375 rpm, which was attributed to their higher efficiency for electron transfer due to their larger surface area. These results indicate that photoelectrochemical efficiency of WO_3 nanoplatelets can be largely enhanced simply by anodizing W under controlled hydrodynamic conditions at a rotation velocity of 375 rpm.

Acknowledgements

Authors would like to express their gratitude for the financial support to the Ministerio de Economía y Competitividad (Project CTQ2013-42494-R) and to the European Social Fund.

References

- [1] H.E. Prakasam, K. Shankar, M. Paulose, O.K. Varghese, C.A. Grimes, *J. Phys. Chem. C* 111 (2007) 7235–7241.
- [2] P. Roy, S. Berger, P. Schmuki, *Angew. Chem. Int. Ed.* 50 (2011) 2904–2939.
- [3] R. Sánchez-Tovar, R.M. Fernández-Domene, D.M. García-García, J. García-Antón, *J. Power Sources* 286 (2015) 224–231.
- [4] H. Tsuchiya, J.M. Macak, I. Sieber, L. Taveira, A. Ghicov, K. Sirotna, P. Schmuki, *Electrochem. Commun.* 7 (2005) 295–298.
- [5] E. Widenkvist, R.A. Quinlan, B.C. Holloway, H. Grennberg, U. Jansson, *Cryst. Growth Des.* 8 (2008) 3750–3753.
- [6] H. Zheng, J.Z. Ou, M.S. Strano, R.B. Kaner, A. Mitchell, K. Kalantar-zadeh, Nanostructured tungsten oxide-properties synthesis, and applications, *Adv. Funct. Mater.* 21 (2011) 2175–2196.
- [7] V. Chakrapani, J. Thangala, M.K. Sunkara, *Int. J. Hydrogen Energy* 34 (2009) 9050–9059.
- [8] V. Cristina, S. Caramori, R. Argazzi, L. Meda, G.L. Marra, C.A. Bignozzi, *Langmuir* 27 (2011) 7276–7284.
- [9] F. Amano, M. Tian, B. Ohtani, A. Chen, *J. Solid State Electrochem.* 16 (2012) 1965–1973.
- [10] Q. Chen, J. Li, B. Zhou, M. Long, H. Chen, Y. Liu, W. Cai, W. Shangguan, *Electrochem. Commun.* 20 (2012) 153–156.
- [11] S.M. Ahmed, J. Leduc, S.F. Haller, *J. Phys. Chem.* 92 (1988) 6655–6660.
- [12] K. Sivila, F. Le Formal, M. Grätzel, *ChemSusChem* 4 (2011) 432–449.
- [13] S.J. Kim, J. Choi, *Electrochem. Commun.* 10 (2008) 175–179.
- [14] J. Zhao, X. Wang, J. Liu, Y. Meng, X. Xu, C. Tang, *Mater. Chem. Phys.* 126 (2011) 555–559.
- [15] T. Bak, J. Nowotny, M. Rekas, C.C. Sorrell, *Int. J. Hydrogen Energy* 27 (2002) 991–1022.
- [16] S. Chen, L.-W. Wang, *Chem. Mater.* 24 (2012) 3659–3666.
- [17] A. Watchareewong, W. Chanmanee, N.R. de Tacconi, C.R. Chenthamarakshan, P. Kajitvichyanukul, K. Rajeshwar, *J. Electroanal. Chem.* 612 (2008) 112–120.
- [18] A.Z. Sadek, H. Zheng, M. Breedon, V. Bansal, S.K. Bhargava, K. Latham, J. Zhu, L. Yu, Z. Hu, P.G. Spizzirri, W. Włodarski, K. Kalantar-zadeh, *Langmuir* 25 (2009) 9545–9551.
- [19] H. Zheng, A.Z. Sadek, K. Latham, K. Kalantar-zadeh, *Electrochem. Commun.* 11 (2009) 768–771.
- [20] K. Kalantar-zadeh, A.Z. Sadek, H. Zheng, V. Bansal, S.K. Bhargava, W. Włodarski, J. Zhu, L. Yu, Z. Hu, *Sens. Actuators B—Chem.* 142 (2009) 230–235.
- [21] W. Lee, D. Kim, K. Lee, P. Roy, P. Schmuki, *Electrochim. Acta* 56 (2010) 828–833.
- [22] Y. Liu, Y. Li, W. Li, S. Han, C. Liu, *Appl. Surf. Sci.* 258 (2012) 5038–5045.
- [23] C.W. Lai, S. Sreekantan, *Mater. Sci. Semicond. Process.* 16 (2013) 303–310.
- [24] C.K. Wang, C.K. Lin, C.L. Wu, S.C. Wang, J.L. Huang, *Electrochim. Acta* 112 (2013) 24–31.
- [25] C. Ng, Y.H. Ng, A. Iwase, R. Amal, *ACS Appl. Mater. Interfaces* 5 (2013) 5269–5275.
- [26] K.R. Reyes-Gil, C. Wiggenhorn, B.S. Brunschwig, N.S. Lewis, *J. Phys. Chem. C* 117 (2013) 14947–14957.
- [27] W. Li, P. Da, Y. Zhang, Y. Wang, X. Lin, X. Gong, G. Zheng, *ACS Nano* 8 (2014) 11770–11777.
- [28] J.H. Ha, P. Muralidharan, D.K. Kim, *J. Alloys Compd.* 475 (2009) 446–451.
- [29] C. Ng, C. Ye, Y.H. Ng, R. Amal, *Cryst. Growth Des.* 10 (2010) 3794–3801.
- [30] J. Yang, W. Li, J. Li, D. Sun, Q. Chen, *J. Mater. Chem.* 22 (2012) 17744–17752.
- [31] S.S. Kalanur, Y.J. Hwang, S.Y. Chae, O.S. Joo, *J. Mater. Chem. A* 1 (2013) 3479–3488.
- [32] M. Anik, T. Cansizoglu, *J. Appl. Electrochem.* 36 (2006) 603–608.
- [33] C.A. Bignozzi, S. Caramori, V. Cristina, R. Argazzi, L. Meda, A. Tacca, *Chem. Soc. Rev.* 42 (2013) 2228–2246.
- [34] J.M. Spurgeon, J.M. Velazquez, M.T. McDowell, *Phys. Chem. Chem. Phys.* 16 (2014) 3623–3631.
- [35] W. Li, J. Li, X. Wang, J. Ma, Q. Chen, *Int. J. Hydrogen Energy* 35 (2010) 13137–13145.
- [36] M. Yagi, S. Maruyama, K. Sone, K. Nagai, T. Norimatsu, *J. Solid State Chem.* 181 (2008) 175–182.
- [37] J.A. Seabold, K.S. Choi, *Chem. Mater.* 23 (2011) 1105–1112.
- [38] R. Solarska, R. Jurczakowski, J. Augustynski, *Nanoscale* 4 (2012) 1553–1556.
- [39] R.M. Fernández-Domene, R. Sánchez-Tovar, E. Segura-Sanchís, J. García-Antón, *Chem. Eng. J.* 286 (2016) 59–67.
- [40] J.Z. Ou, M.Z. Ahmad, K. Latham, K. Kalantar-zadeh, G. Sberveglieri, W. Włodarski, *Procedia Eng.* 25 (2011) 247–251.
- [41] W. Li, J. Li, X. Wang, S. Luo, J. Xiao, Q. Chen, *Electrochim. Acta* 56 (2010) 620–625.
- [42] A. Di Paola, F. Di Quarto, C. Sunseri, *Corros. Sci.* 20 (1980) 1067–1078.
- [43] A.J. Bard, L.R. Faulkner, *Electrochemical Methods: Fundamentals and Applications*, John Wiley & Sons, New York, NY, 2001.
- [44] G.J. Brug, A.L.G. van den Eeden, M. Sluyters-Rehbach, J.H. Sluyters, *J. Electroanal. Chem.* 176 (1984) 275–295.
- [45] B. Hirschorn, M.E. Orazem, B. Tribollet, V. Vivier, I. Frateur, M. Musiani, *Electrochim. Acta* 55 (2010) 6218–6227.
- [46] J.O. Bockris, *J. Appl. Phys.* 52 (1981) 808–810.
- [47] K. Uosaki, H. Kita, *J. Electrochem. Soc.* 130 (1983) 895–897.
- [48] R.M. Fernández-Domene, E. Blasco-Tamarit, D.M. García-García, J. García-Antón, *Electrochim. Acta* 95 (2013) 1–11.
- [49] G. Wang, Y. Ling, H. Wang, X. Yang, C. Wang, J.Z. Zhang, Y. Li, *Energy Environ. Sci.* 5 (2012) 6180–6187.
- [50] C.R. Lhermitte, J.G. Verwer, B.M. Bartlett, *J. Mater. Chem. A* (2016), <http://dx.doi.org/10.1039/C5TA04747A>.
- [51] R. van de Krol, *Principles of Photoelectrochemical Cells*, in: R. van de Krol, M. Grätzel (Eds.), *Photoelectrochemical Hydrogen Production*, New York, 2012.
- [52] A. Hagfeldt, H. Lindström, S. Södergren, S.E. Lindquist, *J. Electroanal. Chem.* 381 (1995) 39–46.
- [53] C. Santato, M. Ulmann, J. Augustynski, *J. Phys. Chem. B* 105 (2001) 936–940.
- [54] J. Chen, D.F. Ollis, W.H. Rulkens, H. Bruning, *Water. Res.* 33 (1999) 669–676.
- [55] M. Ulmann, N.R. de Tacconi, J. Augustynski, *J. Phys. Chem.* 90 (1986) 6523–6530.
- [56] J.C. Hill, K.S. Choi, *J. Phys. Chem. C* 116 (2012) 7612–7620.
- [57] Q. Mi, A. Zhanaidarova, B.S. Brunschwig, H.B. Gray, N.S. Lewis, *Energy Environ. Sci.* 5 (2012) 5694–5700.
- [58] J. Zhao, E. Olide, F.E. Osterloh, *J. Electrochem. Soc.* 162 (2015) H65–H71.
- [59] L. Weinhardt, M. Blum, M. Bär, C. Heske, B. Cole, B. Marsen, E.L. Miller, *J. Phys. Chem. C* 112 (2008) 3078–3082.
- [60] Q. Mi, R.H. Coridan, B.S. Brunschwig, H.B. Gray, N.S. Lewis, *Energy Environ. Sci.* 6 (2013) 2646–2653.
- [61] W.J. Hamer, Y.-C. Wu, *J. Phys. Chem. Ref. Data* 1 (1972) 1047–1075.
- [62] K.S. Pitzer, J.C. Peiperl, R.H. Busey, *J. Phys. Chem. Ref. Data* 13 (1984) 1–102.
- [63] J.M. Prausnitz, R.N. Lichtenhaler, E.G. Azevedo, *Molecular Thermodynamics of Fluid-Phase Equilibria*, Prentice Hall, Upper Saddle River, NJ, 1999.
- [64] M.J. Blandamer, J.B.F.N. Engberts, P.T. Gleeson, J.C.R. Reis, *Chem. Soc. Rev.* 34 (2005) 440–458.
- [65] C. Held, L.F. Cameretti, G. Sadowski, *Fluid Phase Equilib.* 270 (2008) 87–96.
- [66] I.M. Abdulagatov, N.D. Azizov, *Chem. Geol.* 230 (2006) 22–41.
- [67] P.S. Crozier, R.L. Rowley, *Fluid Phase Equilib.* 193 (2002) 53–73.
- [68] G.C. Nihous, C.K. Kinoshita, S.M. Masutani, *Chem. Eng. Sci.* 64 (2009) 2767–2771.
- [69] M. Prego, O. Cabeza, E. Carballo, C.F. Franjo, E. Jiménez, *J. Mol. Liq.* 89 (2000) 233–238.
- [70] J. Chen, D.F. Ollis, W.H. Rulkens, H. Bruning, *Water. Res.* 33 (1999) 661–668.
- [71] T. Kawai, T. Sakata, *J. Chem. Soc. Chem. Commun.* (1980) 694–695.
- [72] W.C. Lin, W.D. Yang, I.L. Huang, T.S. Wu, Z.J. Chung, *Energy Fuel* 23 (2009) 2192–2196.
- [73] A. Galinska, J. Walendziewski, *Energy Fuel* 19 (2005) 1143–1147.